Materials

Glass Frit Filters for Collecting Metal Oxide Nanoparticles

Lyndon B. Johnson Space Center, Houston, Texas

Filter disks made of glass frit have been found to be effective as means of high-throughput collection of metal oxide particles, ranging in size from a few to a few hundred nanometers, produced in gas-phase condensation reactors. In a typical application, a filter is placed downstream of the reactor and a valve is used to regulate the flow of reactor exhaust through the filter. The exhaust stream includes a carrier gas, particles, byproducts, and unreacted particle-precursor gas. The filter selectively traps the particles while allowing the carrier gas, the byproducts, and, in some cases, the unreacted precursor, to flow through unaffected. Although the pores in the filters are much larger than the particles, the particles are nevertheless trapped to a high degree: Anecdotal information from an experiment indicates that 6-nm-diameter particles of MnO_2 were trapped with >99-percent effectiveness by a filtering device comprising a glass-frit disk having pores 70 to 100 μm wide immobilized in an 8-cm-diameter glass tube equipped with a simple twist valve at its downstream end.

This work was done by John Ackerman, Dan Buttry, Geoffrey Irvine, and John Pope of Blue Sky Batteries, Inc., for **Johnson Space Center**. *For further information, contact the Johnson Innovative Partnerships Office at (281) 483-3809.*
MSC-23425

Anhydrous Proton-Conducting Membranes for Fuel Cells

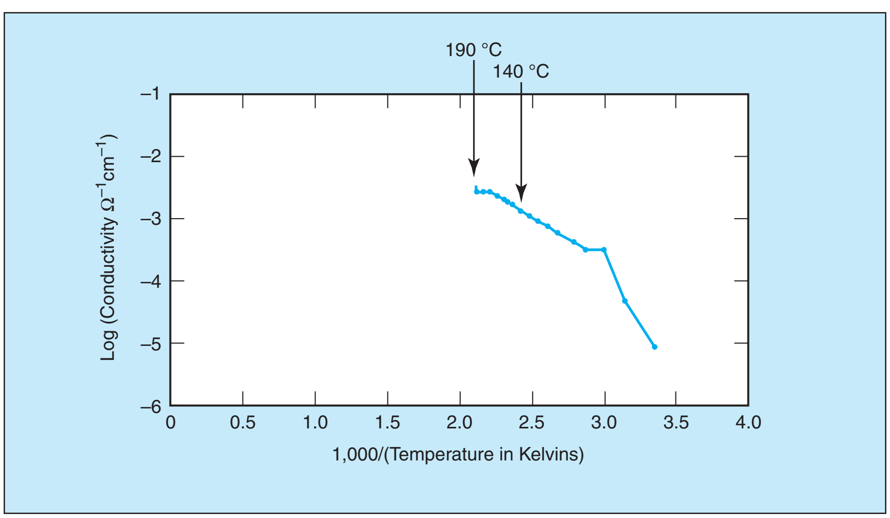
Operating temperatures could be as high as 200 °C.

NASA's Jet Propulsion Laboratory, Pasadena, California

Polymeric electrolyte membranes that do not depend on water for conduction of protons are undergoing development for use in fuel cells. Prior polymeric electrolyte fuel-cell membranes (e.g., those that contain perfluorosulfonic acid) depend on water and must be limited to operation below a temperature of 125 °C because they retain water poorly at higher temperatures. In contrast, the present developmental anhydrous membranes are expected to function well at temperatures up to 200 °C.

The developmental membranes exploit a hopping-and-reorganization proton-conduction process that can occur in the solid state in organic amine salts and is similar to a proton-conduction process in a liquid. This process was studied during the 1970s, but until now, there has been no report of exploiting organic amine salts for proton conduction in fuel cells.

The present development work exploits and extends the previous research on water-free proton conduction in organic amine salts. This work has included an investigation of acid salts of triethylenediamine in which each molecule contains two tertiary nitrogen atoms that can be quaternized. It has been demonstrated that by combining such a proton conductor with nanoparticles of suitable oxide (for example, silica) and a stable binder [for example, poly(tetrafluoroethylene)], one can fabricate a polymeric electrolyte membrane inexpensively. The figure depicts the results of measurements of the ionic conductivity of such a membrane made from triethylenediamine sulfate. The activation energy for proton transport, obtained from the slope of the plot, lies in the range of 0.15 to 0.20 eV — a low range indicative of facile transport of protons.



The **Ionic Conductivity** of a triethylenediamine sulfate membrane was measured as a function of temperature. The conductivity vales are here plotted on a logarithmic scale versus reciprocal of temperature data — a form of plot that facilitates the estimation of activation energy.

Proton-conducting membranes to be investigated in the continuing development effort are divided into the following three classes according to the amine salts and related compounds on which they are based:

- *Type I:* Organic tertiary amine bisulfates, triflates (trifluoromethanesulfomates), and hydrogen phosphates.
- *Type II:* Polymeric quaternized amine bisulfates, triflates, and hydrogen phosphates.
- *Type III:* Polymeric quaternized bisulfates, hydrogen phosphates, and triflates combined with perfluorosulfonic acid-based polymers.

As in the case of the membrane described in the preceding paragraph, a proton-conducting membrane of type I would